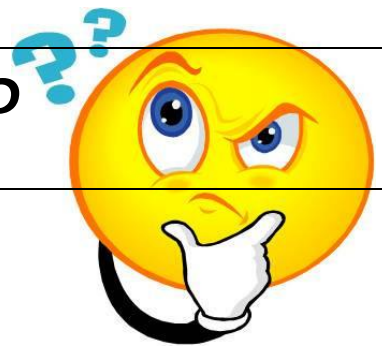


# Is this milk fresh?



## I) Objective :

Lactose is the main carbohydrate in milk. In presence of certain bacteria, there is a lactic fermentation: **Lactose is slowly transformed into lactic acid.**

The measurement of the acidity makes it possible to determine the quantity of lactic acid in the milk. We can deduce its state of freshness and its degree Dornic.

**Definition**: *1°Dornic milk containing 0,1g of lactic acid per liter.*

*The milk is fresh, if its °Dornic is inferior to 18°Dornic (<1,8 g per liter)*

## II) Principle :

The lactic acid of the milk is dosed by a sodium hydroxide solution with colored indicator (phenolphthalein).

When the quantity of soda is insufficient the phenolphthalein will be colorless.

When the quantity will be sufficient, the phenolphthalein becomes pink.

## III) Manipulation :

-Put the sodium hydroxide solution in the graduated burette, adjust to zero;

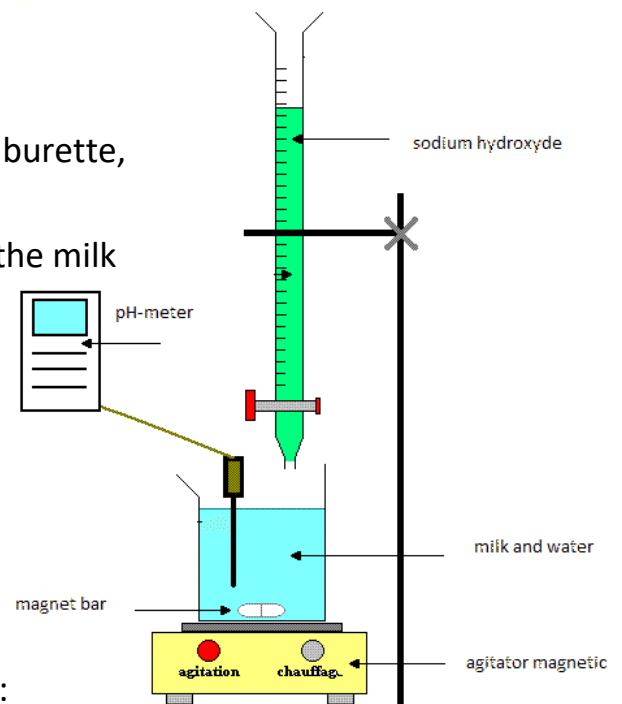
-With a graduated pipette, take 20mL of milk and fill the milk in the beakerflask;

-Add 100mL of distilled water with a graduated Cylinder;

-Add 15 drops of phenolphthalein;

-Place the beakerflask on the magnetic agitator, and shake slowly. Add mL per mL the sodium hydroxide solution, note the pH values.

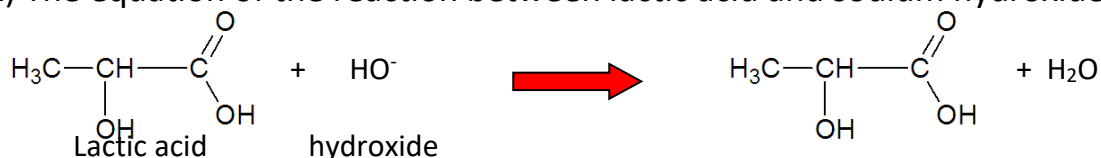
-Write the volume  $V_{HO^-}$  of sodium hydroxide poured :



$V_{HO^-} =$                       mL

#### IV) Results :

1) The equation of the reaction between lactic acid and sodium hydroxide is:



2) The sodium hydroxide solution is  $c = 0.05 \text{ mol.L}^{-1}$ . The quantity of  $\text{OH}^-$  ions contained in the volume  $V_{\text{OH}^-}$  of sodium hydroxide used to dose the milk is deduced from:

$$n(\text{HO}^-) = c \times v_{\text{HO}^-}$$

$n_{\text{HO}^-} =$	mol
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3) The equation shows that there is as much acid as hydroxide:

$n(\text{acidelactique}) = n_{\text{OH}^-} =$	mol <i>in 20 mL of milk</i>
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4) Determine the molar mass  $M$  of lactic acid  $\text{C}_3\text{H}_6\text{O}_3$

Given values:  $M(\text{C}) = 12\text{g.mol}^{-1}$      $M(\text{H}) = 1\text{g.mol}^{-1}$      $M(\text{O}) = 16\text{g.mol}^{-1}$

$M =$	$\text{g.mol}^{-1}$
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Mass of lactic acid **in the 20 mL** of milk:  $m = n \times M =$  g

5) Determine the mass of lactic acid in **1 liter of the milk**.

$m =$	20mL
$m$ in 1 liter	1L=1000ml

$m =$	g of lactic acid in 1L of milk
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Calculate his **Dornic acidity**:

*Is this milk still fit for consumption?*



OR

